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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING The measurement has two significant figures. The measurement has four significant figures. Decimal point is present, so these zeros are significant. The rules are summarized in the following flowchart: qq Nonzero digits 3 4 0 0 g of sulfur pp Nonzero digits 4 0 0 0. mL of oxygen pp

## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 2

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I discovered the Mini-Guide to Problem Solving during my third year of teaching high school chemistry. I have found it invaluable for providing both practice problems and test problems. Each section includes sample problems with step-by-step solutions and a large selection of practice problems at various difficulty levels with the answers given in the back of the book.

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING General Plan for Solving Stoichiometry Problems Convert using the mole ratio A, given in the balanced chemical equation. Mass of substance A Amount in mol of substance A Amount in mol of substance B Convert using the molar mass of A. 1 Mass of substance B 4 2 3 Convert using the molar mass of B. B SAMPLE PROBLEM 1

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## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 8 Stoichiometry

Name \_\_\_\_\_ Date \_\_\_\_\_ Class 2 of 8  
CHEMFILE MINI-GUIDE TO PROBLEM SOLVING  
SAMPLE PROBLEM 1 What is the molarity of a solution that is made by diluting 50.00 mL of a 4.74 M ...

## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 15...

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 1. Determine the percentage composition of each of the following compounds: a. sodium oxalate,  $\text{Na}_2\text{C}_2\text{O}_4$  ans: 34.31% Na, 17.93% C, 47.76% O b. ethanol,  $\text{C}_2\text{H}_5\text{OH}$  ans: 52.13% C, 13.15% H, 34.72% O c. aluminum oxide,  $\text{Al}_2\text{O}_3$  ans: 52.92% Al, 47.08% O d. potassium

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sulfate,  $K_2SO_4$  ans: 44.87% K, 18.40% S, 36.72% O 2.

## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 6

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 14 Concentration of Solutions There are three principal ways to express solution concentration in chemistry—percentage by mass, molarity, and molality. The following table compares these three ways of stating solution concentration. Examining the method of preparation of the three types may

## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 14 ...

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING Neither

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pressure nor temperature is constant; therefore, the combined gas law must be used to calculate the volume of CO<sub>2</sub> at the final temperature and pressure. multiply by the volume ratio, CO<sub>2</sub> use the combined gas law to convert from the initial temperature and pressure to the final temperature and pressure C<sub>2</sub>H<sub>4</sub>

## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 13 ...

Name Date Class CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 18 Titrations

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Empirical Formulas ADDITIONAL PROBLEMS PRACTICE. Name  
Date Class 10 of 11 CHEMFILE MINI-GUIDE TO PROBLEM

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SOLVING 3. Determine the empirical formula for compounds that have the fol-[Filename: HUGPS073.PDF] - Read File Online - Report Abuse

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Name Date Class 4 of 13 CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 3. Aluminum will react with sulfuric acid in the following reaction.  $2\text{Al}(s) + 3\text{H}_2\text{SO}_4(l) \rightarrow \text{Al}_2(\text{SO}_4)_3(aq) + 3\text{H}_2(g)$  a. How many moles of  $\text{H}_2\text{SO}_4$  will react ans: 27 mol  $\text{H}_2\text{SO}_4$  with 18 mol Al? b. How many moles of each product ans: 27 mol H will be produced? 2 9 mol  $\text{Al}_2(\text{SO}_4)_3$  4.

SAMPLE PROBLEM 1 Ammonia is made industrially by reacting ...

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 18 Titrations Chemists have many methods for determining the

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quantity of a substance present in a solution or other mixture. One common method is titration, in which a solution of known concentration reacts with a sample containing the substance of unknown quantity. There are two main requirements

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